

Things to Know, Understand and Do For Chapter 10: Physical Characteristics of Gases

By the end of Chapter 10, you should

Know how to...
State the five points of the kinetic-molecular theory of matter
Define Real and Ideal gases
Describe the following gas properties: expansion, density, fluidity, compressibility, diffusion, and effusion
Tell what conditions would make a gas deviate from ideal behavior
Define pressure and convert between the different units of pressure (mm Hg, torr, atm, and kPa)
Describe how pressure is measured
Perform calculations using Boyle's Law, Charles' Law, Gay-Lussac's Law, Combined Gas Law, and Dalton's Law of Partial Pressures.

understand...
How the kinetic molecular theory affects properties of matter
What makes a gas "ideal" and "real"
That pressure is a measure of collisions of gas molecules with the wall of their containers
How pressure, volume, moles, and temperature of gases are all proportionate to each other according to the kinetic molecular theory

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